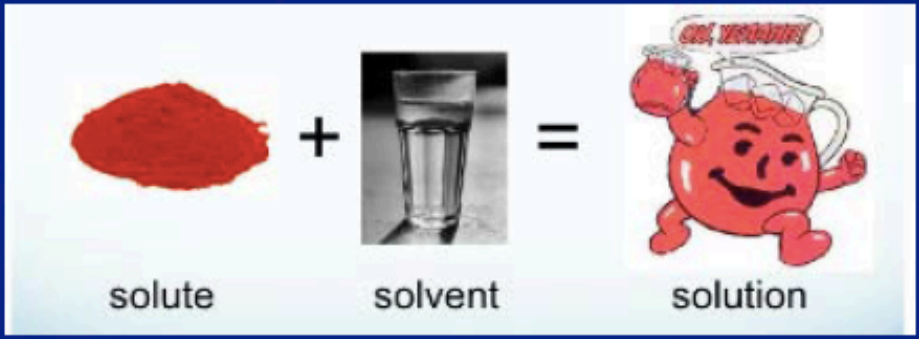
Lab Math
Solutions, Dilutions, Concentrations and Molarity

more substances.

- ***Solute -*** a substance in a solution that is present in the smallest amount.
- ***Solvent -*** a substance in a solution that is present in the largest amount.

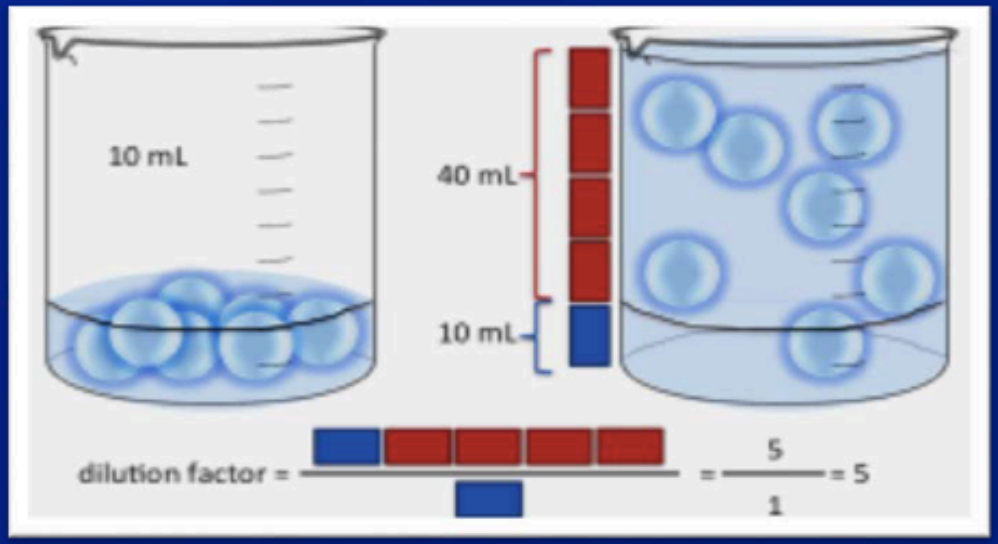


- ***Solubility -*** ability of the solute to dissolve in

Solutions & Dilutions

- **Preparing solutions and making dilutions**
  - Simple dilutions
  - Mixing parts or volumes
  - Serial dilutions
  - Making fixed volumes of specific concentrations from liquid reagents: (***C1)(V1)=(C2)(V2)***
  - Percent solutions (= parts per hundred)
  - Molar solutions (unit=M=moles/L)

A simple dilution is one in which a unit volume of material of interest (solute) is combined with an appropriate volume of a solvent to achieve the desired concentration



✓ The *dilution factor is the total number of unit volumes in which your material will be dissolved.*

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Mixing parts or volumes ≠ simple dilutions

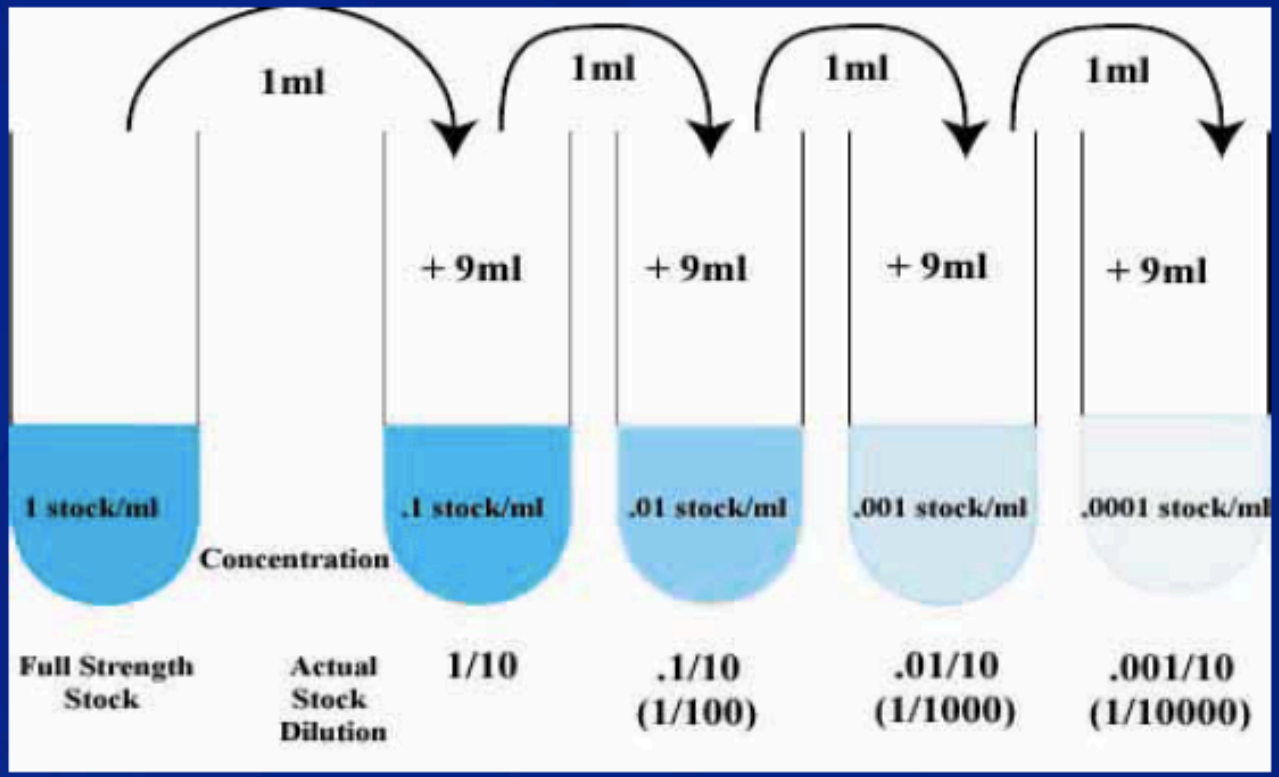
Example: To make up a 1:3 acetic ethanol solution, you are supposed to mix one unit volume of acetic acid and three unit volumes of ethanol.

However, to make 1:3 dilution of acetic acid in ethanol, you would mix one unit volume of acetic acid with two unit volumes of ethanol.

Confused?

"dilution"  "dilute"

A serial dilution is a series of simple dilutions which amplifies the dilution factor quickly. The source of dilution material for each step comes from the diluted material of the previous.



Making fixed volumes of specific concentrations from liquid reagents:

(C1)(V1)=(C2)(V2)

C1V1 -source solution attributes; *C2V2 -new solution attributes*

C= concentration ;* *V= volume**

**In whatever units you are working*

Example: Experiment requires a final concentration of 0.1 M Tris, in a total volume of 2ml vials. You have 1M soln. How much stock solution you will need? (C1)(V1)=(C2)(V2)

(1 M) x V1= (0.1 M) x (2 ml)

V1= (0.1 M) x (2 ml) / (1M) = 0.2ml

Percent solutions (= parts per hundred)

$$\text{Percent by volume (\% (v/v))} = \frac{\text{volume of solute}}{\text{volume of solution}} \times 100\%$$

$$\text{Percent by mass (\% (m/m))} = \frac{\text{mass of solute}}{\text{mass of solution}} \times 100\%$$

Example: A 10% NaCl sol is equal to 10gNaCl in 100ml of solvent, because the solid is measured based upon weight (w) while solvent is measured based upon volume, the NaCl solution should be labeled as 10% w/v NaCl

Molarity (*M*)

- Is a concentration term for solution
- is the number of moles of solute dissolved in one liter of solution.

To calculate the molarity of a solution, divide the moles of solute by the volume of the solution.

$$\text{Molarity } (M) = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Vocabulary to know

Avogadro's Constant: The number of objects in a mole: 6.02×10^{23}

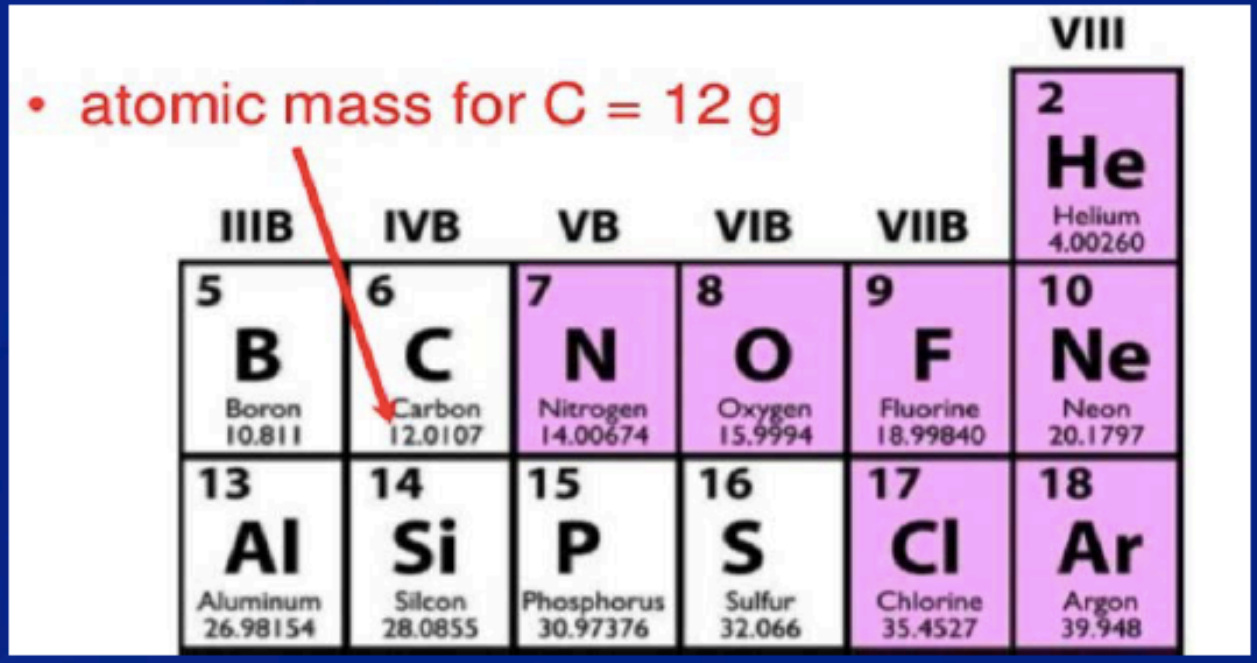
Mole: amount of a substance that contains 6.02×10^{23} particles. A mole of any pure substance has a mass in grams exactly equal to that substance's atomic or molecular mass

Molarity: A unit of concentration equal to the number moles of solute in a 1L of solution.

A mole of any pure substance has a mass in grams exactly equal to that substance's atomic or molecular mass (mass of one mole of a substance).

MM of $CO_2 = {}^{12}C + {}^{16}O_2 = 48$g/mole

(6.02×10^{23} molecules of CO2= 48g)



Molarity dilutions:

- **1M** solution of NaCl (23Na 35Cl) = 58 g NaCl dissolved in 1 liter of H2O
- **1mM** (miliM) = 1:1000 dilution of 1M or 10^{-3}
- **1µM** (microM) = 1:1000 dilution of 1mM or 1:1,000,000 of 1M or 10^{-6}
- **1nM** (nanoM)= 1:1000 dilution of 1 µM or 1:1,000,000,000 of 1M or 10^{-9}
- **1pM** (picoM) = 1:1000 dilution of 1 nM or 1:1,000,000,000,000 of 1M or 10^{-12}

What's the **µM** concentration of a 1:10000 dilution of a solution containing 87 g of NaCl per litter?

$$\text{1M NaCl} = 58\ \text{g/L} \Rightarrow \frac{?M}{87\text{g/L}} = \frac{1M}{58\text{g/L}} \Rightarrow \frac{?M \times 58\cancel{\text{g/L}}}{1M \times 87\cancel{\text{g/L}}} = 1.5M$$

1:10000 dilution of 1.5M = 0.00015 M x 1000 mM = 0.15 mM = 150 **µM**